

Chapter 2**Due: 11:00am on Monday, February 1, 2010****Note:** To understand how points are awarded, read your instructor's [Grading Policy](#).[\[Return to Standard Assignment View\]](#)

H O L

3 4

H

NH₃

H atoms

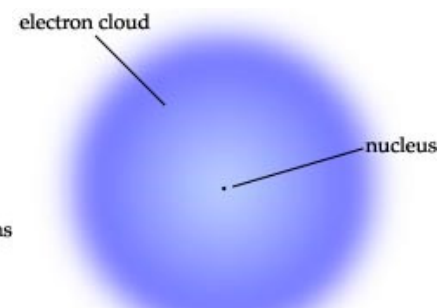
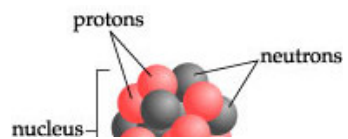
H

NH₃

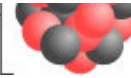
H atoms

1.673×10^{-24} g

1.673×10^{-24} g



nucleus



	g			
	9.109×10^{-28}		-1.602×10^{-19}	-
	1.673×10^{-24}		$+1.602 \times 10^{-19}$	
	1.675×10^{-24}			

 Z A

Fe

H

O

Li

 ${}^7_3\text{Li}$ ${}^7\text{Li}$





$$Z = 14 \quad A = 29$$



 Z A

The screenshot displays a chemistry assignment interface. At the top, there is a row of seven small gray squares. Below this, three panels are arranged horizontally:

- Main group elements:** A periodic table with a highlighted element in the main group.
- Transition metals:** A periodic table with a highlighted element in the transition metal block.
- Inner transition metals:** A periodic table with a highlighted element in the inner transition metal block.

Below the panels, there is a horizontal line.



● ● ● ● ● ● ● ●

Period 1	Period 2	Period 3	Period 4	Period 5	Period 6	Period 7

X

- X amu
- X amu
- X amu
- X amu

X

X

amu

 $^{35}_{17}\text{Cl}$ $^{37}_{17}\text{Cl}$

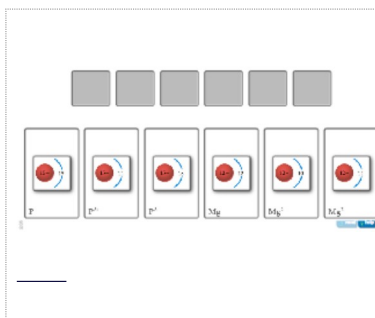
- $^{35}_{17}\text{Cl}$
 $^{37}_{17}\text{Cl}$

 ^{35}Cl

amu

 ^{35}Cl


$$(0.7553)(34.96885) + (0.2447)(36.96590) = 35.45 \text{ amu}$$



<ul style="list-style-type: none"><input type="radio"/><input type="radio"/><input type="radio"/><input checked="" type="radio"/>
<ul style="list-style-type: none"><input checked="" type="radio"/><input type="radio"/><input type="radio"/><input type="radio"/>
H_2O

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••••••••



 Be^{2+}

He

 Al^{3+} Mg^{2+} N^{3-}

Ne

 S^{2-}

Ar

 Se^{2-} Br^{-}

Kr

Ba

CO

CO₂

CO
CO₂

CO₂
CO

- SF₆ g
 - SF_X g
- X

g one-third g

SF₆

X =

CO CO₂

C₂O C₂O₂ CO₂ CO₄

	N	O
	g	g

- NO₂ NO₈ NO₁₀
- N₂O N₂O₄ N₂O₅
- NO NO₂ NO₄
- NO₄ NO₁₀ NO₅

g

g

g

g

g

g

m

g

